Table of Standard Reduction Potentials

Half Reaction	E° (V)
$Cl_{2(g)} + 2\ e^- \rightarrow 2\ Cl^-_{(aq)}$	+1.35827
$Hg^{2+}_{(aq)} + 2\ e^- \rightarrow Hg_{(l)}$	+0.851
$Ag^+_{(aq)} + \ e^- \rightarrow Ag_{(s)}$	+0.7996
$Hg^{2+}_{2(aq)} + 2\ e^- \rightarrow 2\ Hg_{(l)}$	+0.7973
$Fe^{3+}_{(aq)} + \ e^- \rightarrow Fe^{2+}_{(aq)}$	+0.771
$I_{2(s)} + 2\ e^- \rightarrow 2\ I^-_{(aq)}$	+0.5355
$Pb^{2+}_{(aq)} + 2\ e^- \rightarrow Pb_{(s)}$	-0.125
$Sn^{2+}_{(aq)} + 2\ e^- \rightarrow Sn_{(s)}$	-0.137
$Co^{2+}_{(aq)} + 2\ e^- \rightarrow Co_{(s)}$	-0.28
$Cr^{3+}_{(aq)} + \ e^- \rightarrow Cr^{2+}_{(aq)}$	-0.407
$Zn^{2+}_{(aq)} + 2\ e^- \rightarrow Zn_{(s)}$	-0.7618
$Cd(OH)_{2(s)} + 2\ e^- \rightarrow Cd_{(s)} + 2\ OH^-_{(aq)}$	-0.809
$Mn^{2+}_{(aq)} + 2\ e^- \rightarrow Mn_{(s)}$	-1.185